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Oct 28, 2013 References: Lab manual: Koczanski, Krystyna; Xidos, James D. CHEM 1300 Laboratory Manual; UMSU Copy Centre: Presentation On Copper, Cu cthame1. 7,605

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Apr 09, 2008 Best Answer:  $\text{Cu(s)} \rightarrow \text{Cu}^{2+}(\text{aq})$   $E_o = -0.337$  volts  $\text{Cl}_2(\text{g}) + 2\text{Cl}^{-}(\text{aq})$   $E_o = 1.359$  volts \_\_\_\_\_  $\text{Cu(s)} + \text{Cl}_2(\text{g}) + \text{Cu}^{2+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq})$   $E_o = 1.022$  volts

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